Calculate the partial pressure of oxygen in a dry gas at a barometric pressure of 760 Torr and an FiO<sub>2</sub> of 60%?

$$PO_2 = (760 \text{ torr} - 47 \text{ torr}) \times .60$$
  
= 427.8 torr

The patient is breathing room air. Calculate the partial pressure of CO<sub>2</sub> in a dry gas where the barometric pressure is 740 Torr?

The patient is breathing room air. Calculate the partial pressure of  $N_2$  in a dry gas at a barometric pressure of 650 mm Hg?

$$PN_2 = 650 \text{ mm Hg x .78}$$
  
= 507 mm Hg

The patient is breathing room air. The barometric pressure is 750 mm Hg. What is the partial pressure of oxygen after it enters the patient's lungs (saturated gas)? Paco<sub>2</sub> 40 torr & RQ 0.8.

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PAO<sub>2</sub> = [(PBaro - PH<sub>2</sub>O) x .21] - (PaCO<sub>2</sub> / 0.8)
= (703 x .21) - (40/.8)
= 147.6 - 50
= 97.6 mmHg
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## QUESTION:

What is the FiO<sub>2</sub> on top of Pike's Peak at a barometric pressure of 550 mm Hg?

. What is the FiO<sub>2</sub> in a deep, deep well at an atmospheric pressure of 620 mm Hg?

$$FiO_2 = .21$$
 or  $21\% O_2$ 

What is the  $PO_2$  if the PB is 734 Torr and the  $FiO_2$  is 1.0 (100%  $O_2$ )

$$PO_2 = (734 \text{ torr} - 47 \text{ torr}) \times 1$$
  
= 687 torr

## PRACTICE PROBLEM:

A patient being treated for carbon monoxide poisoning is placed in a hyperbaric chamber. The pressure is increased to 3 ATM, the temperature is maintained constant at 37 C. (Note: This is occurring at Leadville Colorado, elevation 10,200 feet). The FiO<sub>2</sub> within the chamber is 1.0 or 100% O<sub>2</sub>. Which of the following statements is true?

I. The inspired oxygen tension ( $PO_2$ ) is 760 mm Hg

II. The inspired oxygen tension is 2280 mm Hg

III. The altitude is not a factor

IV. The altitude and temperature are a factor

V. A hyperbaric chamber is contraindicated for this situation

A. I and IV

B. I, III, IV

C. II and III

D. I, II, III

E. V only